

Unit at a Glance:

September 2015

Monday	Tuesday	Wednesday	Thursday	Friday
	1 First Day of School Class 1 Introduction to Sci10 Start Chem Unit A Pre-test STSK1.3, 2.1, 2.6, 2.8 TextbookA1.1	2 Class 2 Chem History Project STSK 1.1-1.2, 2.7 Textbook A1.2, 1.3 for project-based learning	3 Class 3 Ch1 Quiz STSK 2.2, 2.3, 2.4, 2.8 Textbook A2.0-2.4	4 Class 4 History Project due STSK 3.1-3.4, 3.6, 3.7 TextbookA2.5, 3.0-3.2 Ch 2 Quiz
7 Labour Day – No School	8 Class 5 Practice Problems STSK 3.5 Textbook A3.3	9 Picture Day Class 6 Reactions Lab	10 Class 7 Practice time STSK 3.5 Textbook A3.3 STSK 3.8-3.9 Textbook A3.4	11 Class 8 Ch 3 Quiz Review
14 Class 9 Unit A test Start Physics Unit B				

Day	Outcomes GLO/SLO	Overview	Textbook Reference
Class 1 Part 1	Attitudes: Mutual Respect Safety	<ul style="list-style-type: none"> • Introduction to Science 10 and go over course outline • Get to know names • Student Periodic Table • Careers with chemistry • Pre-tests: Formative assessments of multiple intelligences and learning preferences 	N/A
Part 2	<ul style="list-style-type: none"> • 1.3 careers • Safety attitude 	<ul style="list-style-type: none"> • Start Unit A • The understanding that particles make up the 	Page 6 A 1.0

	<ul style="list-style-type: none"> • 2.1 WHMIS • 2.8 • 2.5 properties • 2.6 solubility 	<p>underlying structure of matter has led to advancements in technology.</p> <ul style="list-style-type: none"> • WHMIS symbols • MSDS • Properties and classification of matter • Physical properties • Chemical properties • Pure substances and mixtures • Mechanical mixture, heterogeneous mixture, homogeneous mixture, suspension, colloids • Give time to work on assignment #1 	A1.1 Safety in the Lab
Class 2 Part 1	<ul style="list-style-type: none"> • 2.7 <p>Project:</p> <ul style="list-style-type: none"> • 1.1 historical • 1.2 atomic model <p>Project Skills:</p> <ul style="list-style-type: none"> • PR • AI • CT 	<ul style="list-style-type: none"> • PROJECT_BASED_LEARNING: • Chemical reactions • Developing ideas about matter • Food chemistry • Heating and freezing, Salting, Fermentation • Metallurgy • Aristotle and matter • Alchemy • Law of conservation of mass • Dalton, Thomson, Rutherford, Bohr, Quantum mechanical model 	Page 10 A 1.2
Part 2	<ul style="list-style-type: none"> • 1.1 historical • 1.2 atomic model 	<ul style="list-style-type: none"> • Project work time • Chapter 1 Review 	Page 18 A 1.3
Class 3 Part 1	<ul style="list-style-type: none"> • 2.2 IUPAC • 2.3 elements combine <ul style="list-style-type: none"> • 2.5 classify ionic and molecular compounds, acids and bases on the basis of their properties; i.e., conductivity, pH, solubility, state 	<ul style="list-style-type: none"> • Chapter 1 Quiz • Assignment #1 due • Elements combine to form many substances, each with its own set of properties. • Periodic table and atomic structure • Elements, metals, nonmetals, metalloids, periods and families, atomic number, atomic mass • Ionic versus molecular 	Page 28 A 2.0 A 2.1

		<ul style="list-style-type: none"> • Compounds • Octet rule • Naming 	
Part 2	<ul style="list-style-type: none"> • 2.4 acids and bases • 2.6 solubility 	<ul style="list-style-type: none"> • Properties and Classification of Ionic and Molecular Compounds • Identifying Ionic Compounds • Special properties of water • Acids and Bases • Acids and bases in the body, properties, indicators • pH scale • Naming acids and bases • Neutralization 	A 2.3 A 2.4
Class 4 Part 1	<ul style="list-style-type: none"> • 3.1 • 3.2 	<ul style="list-style-type: none"> • Project due • Our chemical society • Time to work on Assignment #2 • Chapter 2 Review 	A 2.5
Part 2	<ul style="list-style-type: none"> • 3.3 • 3.4 • 3.6 • 3.7 	<ul style="list-style-type: none"> • Chapter 2 Quiz • Assignment #2 due • Chemical change is a process that involves recombining atoms and energy flows. • Important examples of chemical change • Reactions that form gas or solids • Exothermic and Endothermic • Conservation of mass • Writing Chemical equations • Balancing equations 	A 3.0 A 3.1 A 3.2
Class 5 Part 1		<ul style="list-style-type: none"> • Practice problems • Time to work on Assignment #3 	A3
Part 2	<ul style="list-style-type: none"> • 3.5 	<ul style="list-style-type: none"> • 5 types of chemical reactions: • formation • decomposition • Predicting products of chemical reactions 	A 3.3

		<ul style="list-style-type: none"> • Time to work on Assignment #3 	
Class 6 Part 1	<ul style="list-style-type: none"> • 3.5 	<ul style="list-style-type: none"> • 5 types of chemical reactions: • hydrocarbon • single replacement • double replacement • Predicting products of chemical reactions • Time to work on Assignment #3 	A 3.3
Part 2	Skills Initiating and Planning Performing and Recording Analyzing and interpreting Communication and teamwork	<ul style="list-style-type: none"> • Five types of chemical reactions lab 	LAB
Class 7 Part 1	Same as before	<ul style="list-style-type: none"> • Finish Lab write up • 5 types of chemical reactions • Time to work on Assignment #3 	A 3.3
Part 2	<ul style="list-style-type: none"> • 3.8 • 3.9 	<ul style="list-style-type: none"> • Moles • Avagadro's number • Calculating molar mass • Moles and the law of conservation of mass • Time for Assignment #3 • Chapter 3 review 	Page 107 A 3.4
Class 8 Part 1		<ul style="list-style-type: none"> • Chapter 3 Quiz • Optional lab or time to fill with extra materials 	LAB
Part 2		<ul style="list-style-type: none"> • Finish Lab (or this is time that can be spent going over the more difficult concepts that students are having difficulties with) • Review 	LAB Review
Class 9 Part 1	Unit A test	TEST	TEST
Part 2	Start next unit		